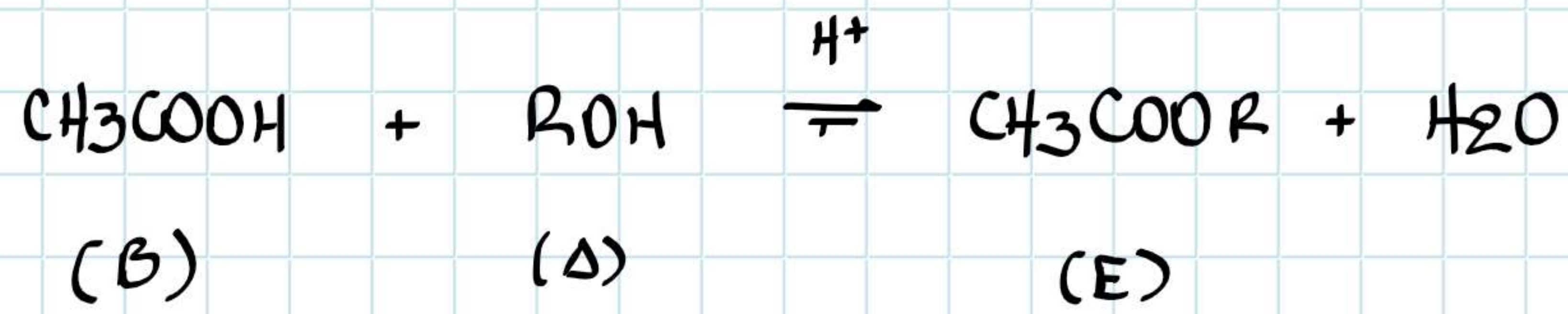
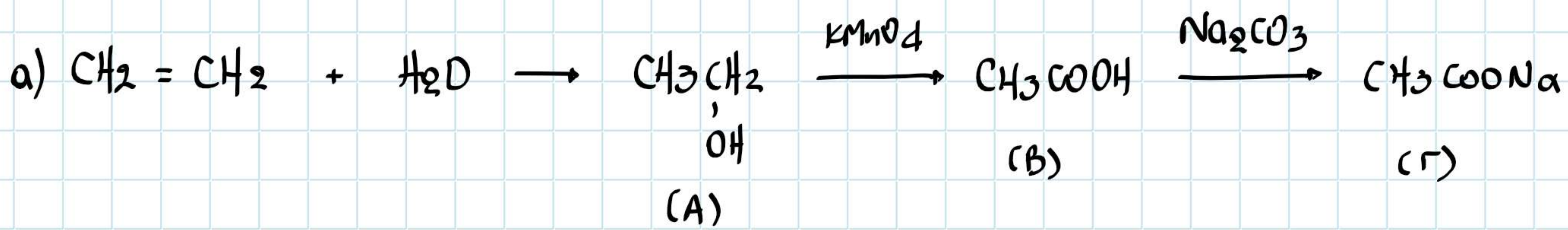
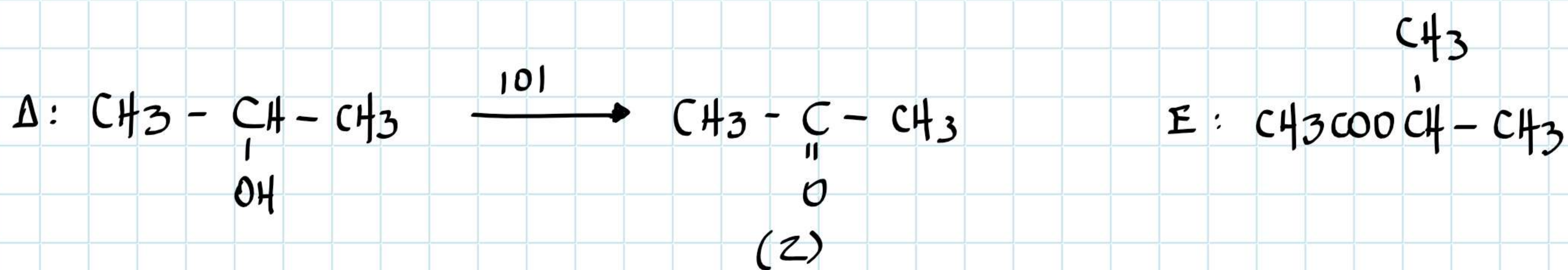
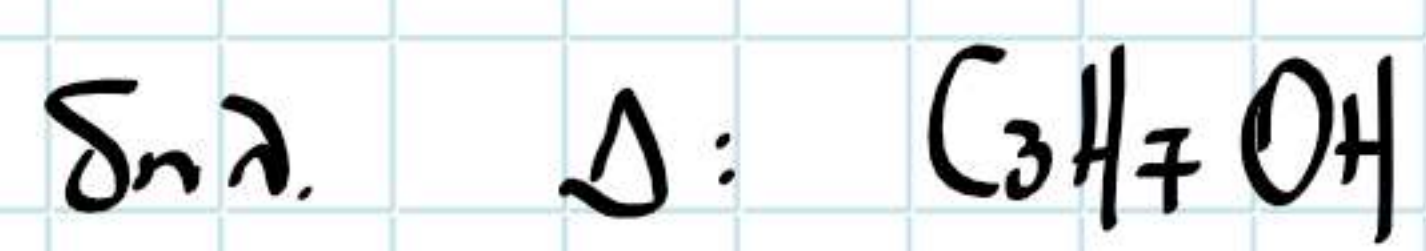


Άσκηση 4.13.



$$M_{rE} = 102 \Rightarrow 12 + 3 + 12 + 2 \cdot 16 + 12v + 2u + 1 = 102$$

$$\Rightarrow v = 3$$

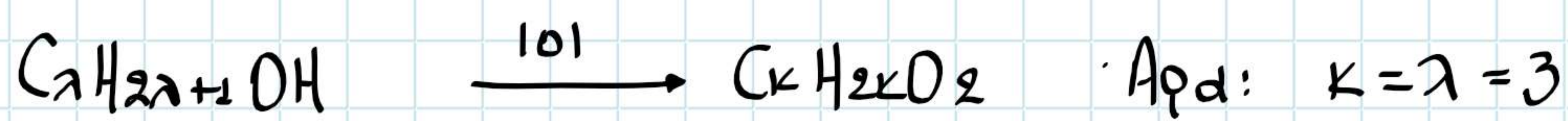
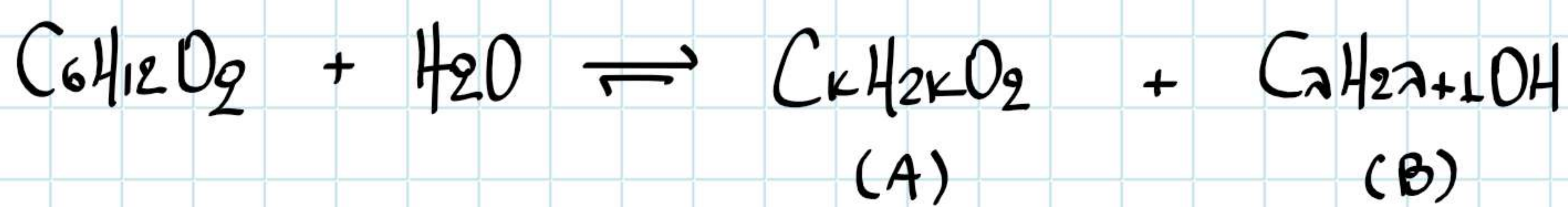


β) Α, Β και το Δ

Άσκηση 4.14.

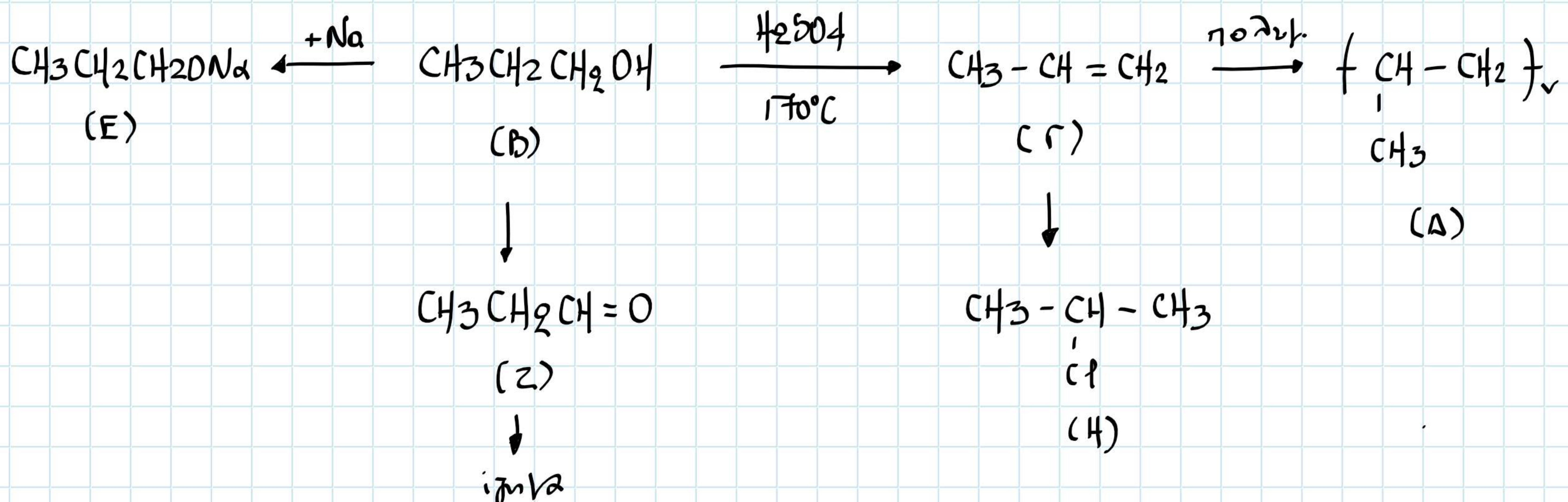
Εστέρας:  $\text{C}_x\text{H}_{2x}\text{O}_2$

$$\frac{m_c}{m_o} = \frac{9}{4} \Rightarrow \frac{12v}{32} = \frac{9}{4} \Rightarrow 48v = 9 \cdot 32 \Rightarrow \underline{v=6}$$

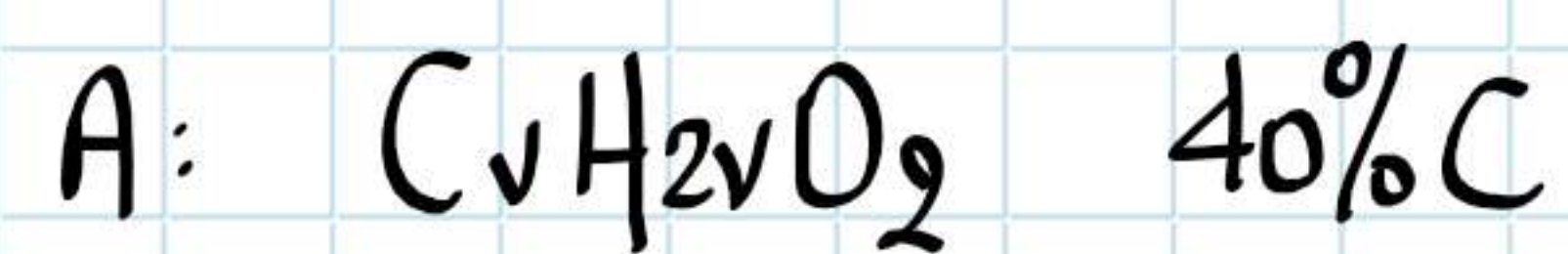


Οπότε: Α:  $\text{CH}_3\text{CH}_2\text{COOH}$

Β:  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$



### Άσκηση 4.15



α) Σε 100g  $C_vH_{2v}O_2$  έχουμε 40g C

$$\frac{14v+32}{100} = \frac{40}{100}$$

$$100 \cdot 12v = 40 \cdot (14v + 32) \Rightarrow 1200v = 560v + 1280$$

$$\Rightarrow 640v = 1280 \Rightarrow \underline{v=2}$$

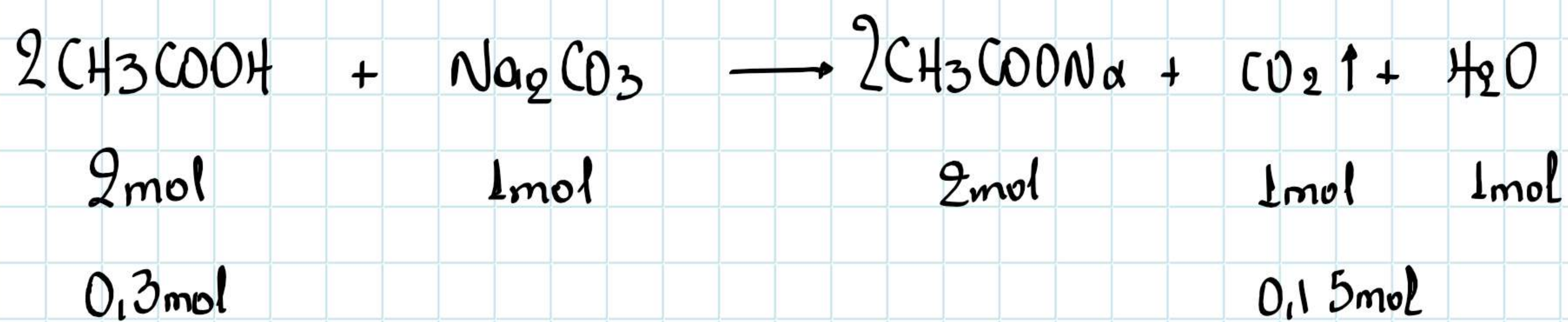


Καρβ. οξύ:  $CH_3COOH$  : αιθανικό οξύ

Εστέρας:  $CH_3COOCH_3$  : μεθανικός μεθυλεστέρας

β) 18g  $CH_3COOH$   $M_r=60 \Rightarrow n = \frac{18}{60} = 0,3 \text{ mol}$

Η ένωση Α είναι οξύ ηλεϊδί αντιδρά με  $Na_2CO_3$ .



$$V_{CO_2} = 0,15 \cdot 22,4 = 3,36 \text{ L}$$

### Άσκηση 4.16

Έστω  $x \text{ mol}$   $CH_3CH_2OH$  και  $y \text{ mol}$   $CH_3CH=O$

$$m_{\mu} = 18 \Rightarrow m_1 + m_2 = 18 \Rightarrow x \cdot 46 + y \cdot 44 = 18 \quad (1)$$

Οξειδώσεις:

