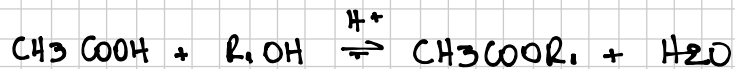


22.18.

$$36 \text{ g CH}_3\text{COOH } M_r=60 \Rightarrow n = \frac{36}{60} = 0,6 \text{ mol}$$

a)



опх: 0,6 mol 0,6 mol

авт/нар: -w -w +w +w

хул.лоор: 0,6-w 0,6-w w w

$$K_c = \frac{\left(\frac{w}{v}\right)^2}{\left(\frac{0,6-w}{v}\right)^2} \Rightarrow 2 = \frac{w}{0,6-w} \Rightarrow 1,2 - 2w = w \Rightarrow \boxed{w = 0,4} \text{ mol}$$

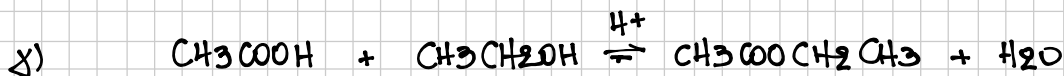
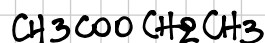
$$a = \frac{0,4}{0,6} = 0,67 \text{ и } 67\%$$

b)

$$0,4 = \frac{35,2}{M_r} \Rightarrow M_r = 88 \Rightarrow 27 + 32 + 14v + 1 = 88$$

$$14v = 28$$

$$\boxed{v = 2}$$



опх: 0,6 mol (0,6 + k)

авт: -x -x x x

хул.лоор: 0,6-x 0,6+k-x x x

$$a = \frac{x}{0,6} = 0,8 \Rightarrow x = 0,48$$

$$4 = \frac{0,48 \cdot 0,48}{0,12 \cdot (0,12 + k)} \Rightarrow 0,12 + k = 0,48 \Rightarrow \underline{\underline{k = 0,36}}$$