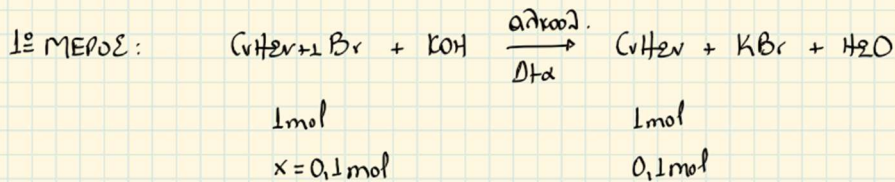
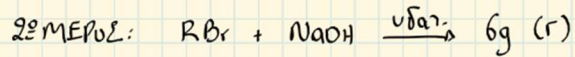
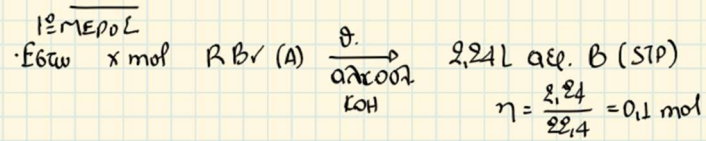
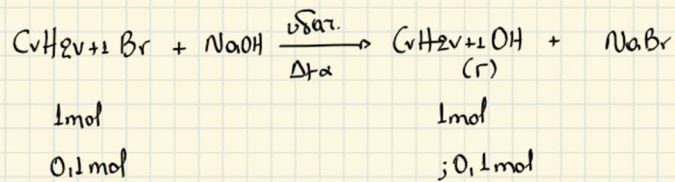


Άσκηση 2213

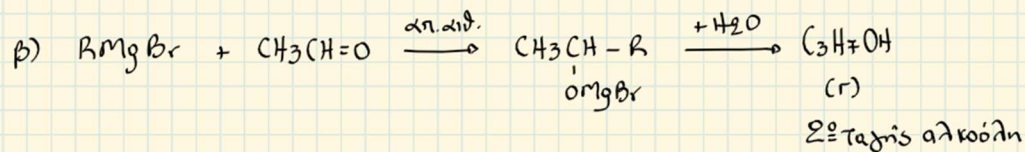


2^ο ΜΕΡΟΣ: Έχουμε $0,1 \text{ mol}$ $C_nH_{2n+1}Br$

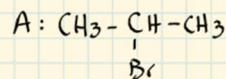
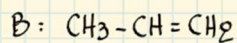
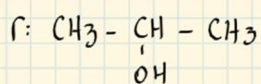


$$n_r = \frac{m}{M_r} \Rightarrow 0,1 = \frac{6}{M_r} \Rightarrow M_r = 60 \Rightarrow 14n + 18 = 60 \Rightarrow \underline{n=3}$$

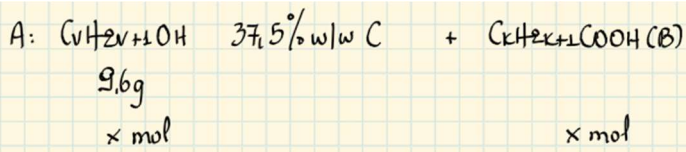
a) Μ.Τ. A: $C_3H_7Br \Rightarrow 0,2 \text{ mol}$, $M_r = 123 \Rightarrow m = 0,2 \cdot 123 = 24,6 \text{ g}$.



Αρα: $R \Rightarrow CH_3-$



22.14



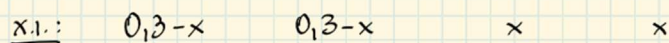
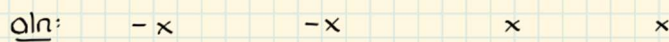
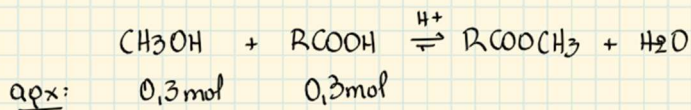
στη x.i. 17,6g εστέρας. $k_c = 4$

Στα 100g ένωσης περιέχονται 37,5g C

Στα $14v+18$ $12v$ g

$$\text{οπότε: } 100 \cdot 12v = 37,5 \cdot (14v+18) \Rightarrow \underline{v=1} \quad A: CH_3OH \quad M_r = 32$$

$$\eta_A = \frac{9,6}{32} = 0,3 \text{ mol}$$



$$K_c = \frac{4 \left(\frac{x}{v}\right)^2}{\left(\frac{0,3-x}{v}\right)^2} \Rightarrow 2 = \frac{x}{0,3-x} \Rightarrow 0,6 - 2x = x \Rightarrow \underline{x=0,2}$$

$$\eta_{\text{εστ.}} = \frac{m}{M_r} \Rightarrow 0,2 = \frac{17,6}{M_r} \Rightarrow M_r = 88$$

$$14k+1+12+32+12+3=88$$

$$k=2$$

